

Participation Assignment

CHEM 1050-Chemistry and the Citizen

Name:

#3

Section: 03, MW

Due Date: Wednesday 1/13/2016

1. 57.152 g of solid gold at 75.3 °C is plunged into a container of cold water. Once the water and the gold have attained the same temperature, it's determined that 87.6 J of heat have been transferred from the gold sample to the water. What is the new temperature of the solid gold? The specific heat of gold is 0.129 J/g °C.

2. List the charge and the mass of each of the following subatomic particles:

a. proton

b. neutron

c. electron

3. List the numbers of protons, electrons, and neutrons in each of the following isotopes:

a. carbon-12

b. carbon-13

c. carbon-14

3. Write the isotopic symbol for each of the following isotopes:

a. carbon-12

b. carbon-13

c. carbon-14