

Preclass Assignment

CHEM 1090-General Chemistry I

Name:

#9

Section: 31, MWF

Due Date: Wednesday 11/1/2017

1. How many grams of KCl are needed to make 500.0 mL of a 0.500 M solution?
2. Using the “Aqueous Solubility Rules for Ionic Compounds” handout, state whether each of the following compounds are soluble or insoluble:
 - a. sodium chloride
 - b. sodium nitrate
 - c. sodium nitrite
 - d. sodium hydroxide
 - e. silver chloride
 - f. lead(II) chloride
 - g. silver nitrate
 - h. iron(III) hydroxide
3. When sodium hydroxide is dissolved in water, what ions are formed? Write both the name and chemical formula for both the cation and the anion. Hint: why is it named “sodium hydroxide”?
4. When solutions of sodium hydroxide and iron(III) nitrate are mixed together, do you expect a precipitate to be formed? If so, write both the chemical formula and the name of the precipitate.
5. Write the chemical formula for each of the acids that are commonly listed as strong acids.