

# Participation Assignment

## CHEM 1090-General Chemistry I

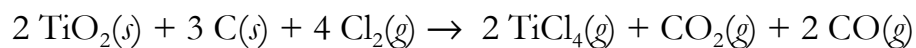
Name:

#9

Section: 31, MWF

Due Date: Wednesday 8/2/2017

1a. Calculate the mass in grams of titanium(IV) chloride,  $\text{TiCl}_4$ , produced when  $9.70 \times 10^8$  g  $\text{TiO}_2$ , are reacted with an excess of carbon and chlorine:



1b. If  $2.17 \times 10^9$  g  $\text{TiCl}_4(g)$  are actually produced, what is the percent yield?