

Participation Assignment

CHEM 1090-General Chemistry I

Name:

#9

Section: 32, TR

Due Date: Thursday 11/1/2018

1. Using the “Aqueous Solubility Rules for Ionic Compounds” handout, state whether each of the following compounds are soluble or insoluble:

a. sodium chloride

e. silver chloride

b. sodium nitrate

f. lead(II) chloride

c. sodium nitrite

g. silver nitrate

d. sodium hydroxide

h. iron(III) hydroxide

2. When sodium hydroxide is dissolved in water, what ions are formed? Write both the name and chemical formula for both the cation and the anion. Hint: why is it named “sodium hydroxide”?

3. When solutions of sodium hydroxide and iron(III) nitrate are mixed together, do you expect a precipitate to be formed? If so, write both the chemical formula and the name of the precipitate.

4. Write the chemical formula for each of the acids that are commonly listed as strong acids.