

# Preclass Assignment

## CHEM 1090-General Chemistry I

Name:

#6

Section: 32, TR

Due Date: Tuesday 11/6/2018

1. Categorize each of the following as either strong, weak, or nonelectrolytes:

a. NaCl

e. BaCl<sub>2</sub>

b. HCl

f. HC<sub>2</sub>H<sub>3</sub>O<sub>2</sub>

c. H<sub>2</sub>O (a molecule)

g. H<sub>2</sub>SO<sub>4</sub>

d. AgNO<sub>3</sub>

h. Ca(NO<sub>3</sub>)<sub>2</sub>

2. Write molecular, total ionic, and net ionic equations for the reaction between silver nitrate and sodium chloride.

3. How many grams of KCl are needed to make 500.0 mL of a 0.500 M solution?

4. Name each of the following acids (look in section 2.7):

a. HCl

d. HNO<sub>2</sub>

b. HBr

e. H<sub>2</sub>SO<sub>4</sub>

c. HNO<sub>3</sub>

f. H<sub>3</sub>PO<sub>4</sub>