

Participation Assignment

CHEM 1090-General Chemistry I

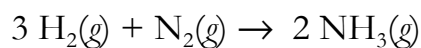
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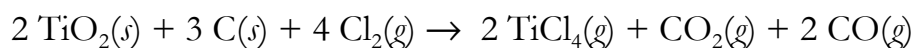
Section: 33, TR

Due Date: Tuesday 2/5/2019

1. Calculate the mass of NH_3 that can be produced from 10.00 g H_2 and 40.00 g N_2 . Also, identify the limiting reactant.



2a. Calculate the mass in grams of titanium(IV) chloride, TiCl_4 , produced when 9.70×10^8 g TiO_2 , are reacted with an excess of carbon and chlorine:



2b. If 2.17×10^9 g $\text{TiCl}_4(g)$ are actually produced, what is the percent yield?