

Preclass Assignment

CHEM 1090-General Chemistry I

Name:

#6

Section: 32, TR

Due Date: Thursday 11/2/2017

1. Categorize each of the following as either strong, weak, or nonelectrolytes:

a. NaCl

e. BaCl₂

b. HCl

f. HC₂H₃O₂

c. H₂O (a molecule)

g. H₂SO₄

d. AgNO₃

h. Ca(NO₃)₂

2. Write molecular, total ionic, and net ionic equations for the reaction between silver nitrate and sodium chloride.

3. How many grams of KCl are needed to make 500.0 mL of a 0.500 M solution?

4. Name each of the following acids (look in section 2.7):

a. HCl

d. HNO₂

b. HBr

e. H₂SO₄

c. HNO₃

f. H₃PO₄