Participation Assignment
CHEM 1100-General Chemistry II

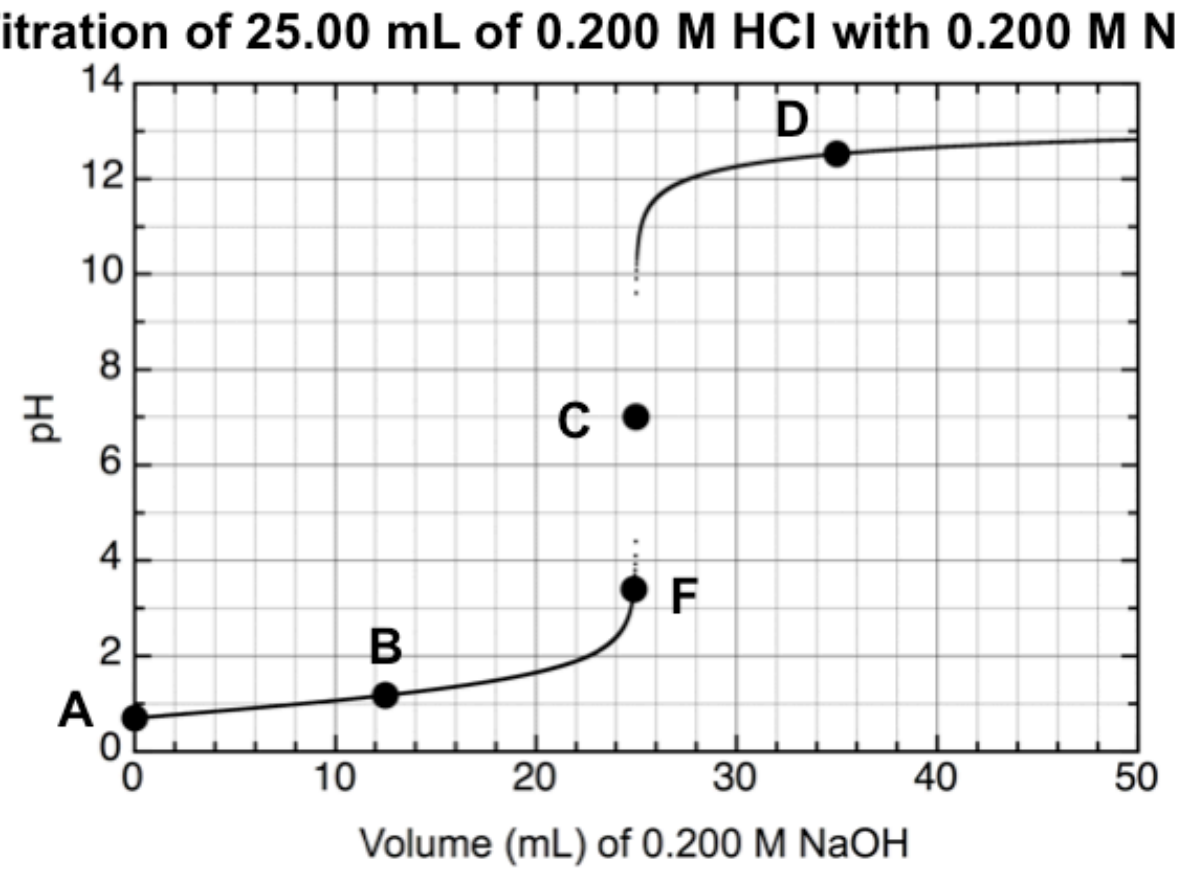
Name: **#12**

Section: 31, TR **Due Date: Tuesday 2/19/2019**

1. State whether you expect solutions of each of the following substances to be acidic, basic, or neutral:

a. HCl b. NaOH c. N_2H_4 d. NaCl e. $AlCl_3$

f. NaClO g. NH_4NO_3 h. NH_4F i. $NaHCO_3$



2. The above graph represents the titration of 25.00 mL 0.200 mol/L HCl by a solution that is 0.200 mol/L NaOH.

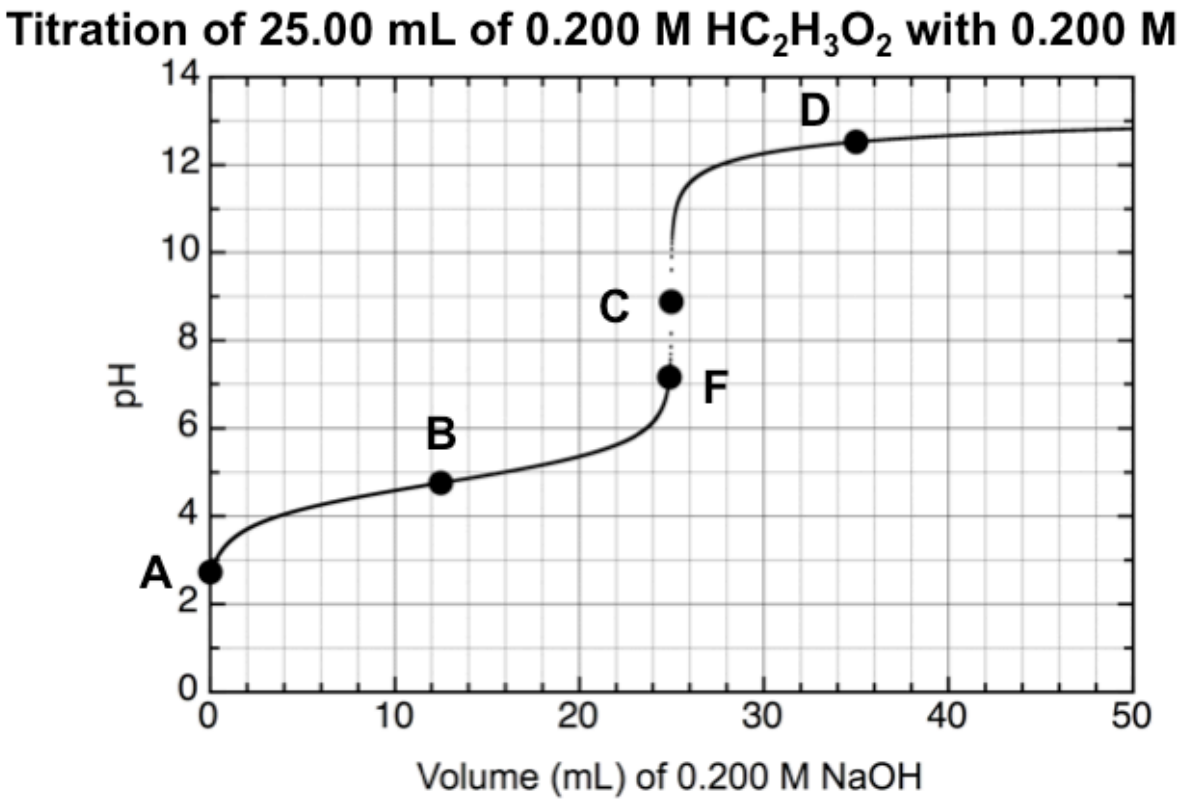
Point A: Calculate the pH before the addition of any NaOH.

Point B: Calculate the pH after the addition of 12.50 mL of the NaOH solution.

Point C: Calculate the pH after the addition of 25.00 mL of the NaOH solution.

Point D: Calculate the pH after the addition of 35.00 mL of the NaOH solution.

3. Define "equivalence point".



4. The above graph represents the titration of 25.00 mL 0.200 mol/L $HC_2H_3O_2$ by a solution that is 0.200 mol/L NaOH.

Point A: Calculate the pH before the addition of any NaOH.