

Participation Assignment

CHEM 1100-General Chemistry II

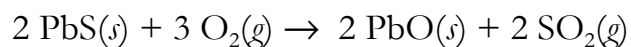
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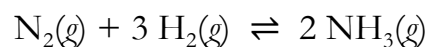
Section: , TR

Due Date: Tuesday 3/5/2019

1. Calculate the temperature at which the following reaction becomes nonspontaneous:



2. Calculate the free energy of reaction for the following chemical reaction at 25 °C under the listed conditions:



$$P(\text{N}_2) = 100.0 \text{ atm}$$

$$P(\text{H}_2) = 40.0 \text{ atm}$$

$$P(\text{NH}_3) = 60.0 \text{ atm}$$

3. Calculate the equilibrium constant for the following chemical reaction at 25 °C:

