

Participation Assignment

CHEM 1100-General Chemistry II

Name:

#10

Section: 31, TR

Due Date: Tuesday 5/1/2018

1. Define each of the following:

a. Arrhenius Acid

b. Arrhenius Base

c. Brønsted-Lowry Acid

d. Brønsted-Lowry Base

2. Fill in the following table. See section 16.3 in your textbook. I will be checking for the proper number of significant figures (see p716 for the proper number when writing the pH). The book doesn't discuss pOH (I will cover it in class) but you can determine it by using:
$$\text{pH} + \text{pOH} = 14.00$$

$[\text{H}_3\text{O}^{1+}]$	$[\text{OH}^{1-}]$	pH	pOH
1.0×10^{-7}			
1.0×10^{-5}			
2.4×10^{-8}			
		10.28	
			12.89

3. Calculate the pH of a solution that is 0.025 M HCl.

4. Calculate the pH of a solution that is 0.025 M $\text{HC}_2\text{H}_3\text{O}_2$.