

# Participation Assignment

## CHEM 1100-General Chemistry II

Name:

#14

Section: 31, TR

Due Date: Tuesday 5/15/2018

1. The solubility of silver chloride, AgCl, is  $1.9 \times 10^{-6}$  g/mL at 25 °C. Calculate the  $K_{sp}$  of silver chloride.

2. Calculate the molar solubility of  $MgF_2$ .

3. Calculate the molar solubility of  $MgF_2$  in 0.025 M NaF. The  $K_{sp}$  for  $MgF_2$  is  $5.16 \times 10^{-11}$ .

4. Will a precipitate form if equal volumes of aqueous solutions of  $1.0 \times 10^{-5}$  mol/L  $\text{Ba}(\text{NO}_3)_2$  and  $4.0 \times 10^{-5}$  mol/L  $\text{Na}_2\text{CO}_3$  are mixed?