

Standard Enthalpies of Formation



| Substance | ΔH°_f (kJ/mol) | Substance | ΔH°_f (kJ/mol) | Substance | ΔH°_f (kJ/mol) |
|---|-----------------------------|---|-----------------------------|-------------------------------------|-----------------------------|
| AgCl(s) | -127.0 | CaSO ₄ (s) | -1434.5 | N ₂ H ₄ (g) | +95.4 |
| Al ₂ O ₃ (s) | -1675.7 | Fe ₂ O ₃ (s) | -824.2 | N ₂ H ₄ (l) | +50.6 |
| CHCl ₃ (g) | -103.2 | HBr(g) | -36.4 | N ₂ O(g) | +82.1 |
| CH ₂ Cl ₂ (g) | -95.5 | HCl(g) | -92.3 | N ₂ O ₄ (g) | +9.1 |
| CH ₂ O(g) | -115.9 | HF(g) | -272.6 | N ₂ O ₅ (g) | +11.3 |
| CH ₃ Cl(g) | -83.7 | HC ₂ H ₃ O ₂ (l) | -483.5 | NaCl(s) | -411.1 |
| CH ₃ OH(l) | -238.4 | H ₂ O(g) | -241.8 | NaHCO ₃ (s) | -950.8 |
| CH ₄ (g) | -74.9 | H ₂ O(l) | -285.8 | NaOH(s) | -425.8 |
| CO(g) | -110.5 | H ₂ O ₂ (l) | -187.8 | Na ₂ CO ₃ (s) | -1130.8 |
| CO ₂ (g) | -393.5 | H ₂ S(g) | -20.5 | O(g) | +249.2 |
| C ₂ H ₂ (g) | +226.7 | H ₂ SO ₄ (l) | -814.0 | O ₃ (g) | +142.7 |
| C ₂ H ₄ (g) | +52.5 | KCl(s) | -436.7 | PCl ₃ (g) | -288.7 |
| C ₂ H ₆ (g) | -84.0 | K ₂ SO ₄ (s) | -1437.7 | PCl ₅ (g) | -374.9 |
| C ₂ H ₅ OH(l) | -276.0 | MgCl ₂ (s) | -641.6 | PbO(s) | -219.4 |
| C ₃ H ₅ N ₃ O ₉ (l) | -370.0 | MgO(s) | -601.2 | PbS(s) | -98.3 |
| C ₆ H ₆ (l) | +49.0 | MgCO ₃ (s) | -1111.7 | SF ₆ (g) | -1220.5 |
| CaCO ₃ (s) ¹ | -1207.6 | NH ₃ (g) | -45.9 | SO ₂ (g) | -296.8 |
| CaCl ₂ (s) | -795.8 | NH ₄ Cl(s) | -314.6 | SO ₃ (g) | -395.8 |
| CaO(s) | -635.1 | NO(g) | +90.3 | | |
| Ca(OH) ₂ (s) | -986.1 | NO ₂ (g) | +33.1 | | |

¹calcite

All standard state, 25 °C and 1 bar (written to 1 decimal place).

P.J. Linstrom and W.G. Mallard, Eds, NIST Chemistry WebBook, NIST Standard Reference Database Number 69, National Institute of Standards and Technology, Gaithersburg MD, 20899, <http://webbook.nist.gov>, (retrieved March 9, 2011).

Lide, David R., Ed., Handbook of Chemistry and Physics, 84th ed., CRC Press: Boca Raton FL, 2003, 5-1 to 5-60.